## Information: Weighted Averages

Examine the table of student test scores for five tests they have taken.

| Test | Student A Student B |  |
| :--- | :--- | :---: |
| 1 | 95 | 76 |
| 2 | 74 | 88 |
| 3 | 82 | 90 |
| 4 | 92 | 81 |
| 5 | 81 | 72 |
| Average Grade: | - | - |

## Critical Thinking Questions

1. Calculate the average grade for students A and B and enter the average in the table above.
2. If you know a student's average grade can you tell what the student's individual test scores were? Explain.
3. Suppose student C had an average of $83 \%$. On each of his five tests he scored either $65 \%$ or $95 \%$. Which score occurred more often? Explain.
4. What if the teacher decided that test five would count for $40 \%$ of the final grade and test four would count for $30 \%$ of the final grade and each of the other tests would count for $10 \%$.
Calculate the new average for each student. Note: this is called the weighted average.
Student A's new average: $\qquad$ Student B's new average: $\qquad$

## Critical Thinking Questions

5. Neon has three different isotopes. $90.51 \%$ of neon atoms have a mass of $19.992 \mathrm{amu} .0 .27 \%$ of neon atoms have a mass of $20.994 \mathrm{amu} .9 .22 \%$ of neon atoms have a mass of 21.991 amu . What is the average atomic mass of neon?
6. Chlorine- 35 is one isotope of chlorine. ( 35 is the mass number.) Chlorine- 37 is another isotope of chlorine. How many protons and how many neutrons are in each isotope of chlorine?
7. Of all chlorine atoms, $75.771 \%$ are chlorine- 35 . Chlorine- 35 atoms have a mass of 34.96885 amu. All other chlorine atoms are chlorine-37 and these have a mass of 36.96590. Calculate the average atomic mass of chlorine.
8. Do your answers for questions 5 and 7 agree with the average atomic masses for neon and chlorine on the periodic table?
9. A certain element has two isotopes. One isotope, which has an abundance of $72.15 \%$ has a mass of 84.9118 amu. The other has a mass of 86.9092 amu . Calculate the average atomic mass for this element. 10. Given the following data, calculate the average atomic mass of magnesium.

| Isotope | Mass of Isotope | Abundance |
| :--- | :--- | :--- |
| Magnesium-24 | 23.985 amu | $78.70 \%$ |
| Magnesium- 25 | 24.986 amu | $10.13 \%$ |
| Magnesium- 26 | 25.983 amu | $11.17 \%$ |

