

Practice Problems (1-step conversions using Avogadro's number)

Modern Chemistry, p 86

Instructions: Complete the conversions indicated below. You MUST use dimensional analysis (conversion factors) and you MUST show all units. Sig figs count!!!

1. How many moles of lead, Pb, are in 1.50×10^{12} atoms of lead?
2. How many moles of tin, Sn, are in 2500 atoms of tin?
3. How many atoms of aluminum, Al, are in 2.75 mol of aluminum?