CHEMISTRY I, IH

Class Work 12-3

NAME _____

1. COMPLETE THE EQUATION, IF NECESSARY
2. BALANCE THE EQUATION.
3. SOLVE FOR THE UNKNOWN

PERIOD _____

GIVEN			UNKNOWN
0.25 mol HgO	1.	$2HgO \rightarrow 2Hg + 1O_2$	How many moles of Hg?
0.50 mol NaCl	2.	$2NaCl \rightarrow 2Na + 1Cl_2$	How many moles of Cl_2?
2.0 g H_2	3.	$1N_2 + 3H_2 \rightarrow 2NH_3$	How many grams of NH_3?
1.25 g CuO	4.	$2CuO \rightarrow 2Cu + 1O2$	How many grams of O_2?
2.50 mol Li	5.	$6Li + 1N_2 \rightarrow 2Li_3N$	How many moles of Li_3N?
0.33 mol Ag_2S	6.	$16Ag + 1S_8 \rightarrow 8Ag_2S$	How many moles of Ag?
0.67 g Na_2O	7.	$2NaOH \rightarrow 1Na_2O + 1H_2O$	How many grams of NaOH?
0.125 g NaOH	8.	$2Na + 2HOH \rightarrow 2NaOH + 1H_2$	How many grams of HOH?
0.375 mol Na_2SO_4	9.	$2NaCl + H_2SO_4 \rightarrow 2HCl + 1Na_2SO_4$	How many moles of NaCl?
10.0 g Al_2O_3	10.	$2Al + 1Fe_2O_3 \rightarrow 1Al_2O_3 + 2Fe$	How many moles of Al?
25.5 g $Ca(OH)_2$	11.	$1Ca(OH)_2 + 1Mg(HCO_3)_2 \rightarrow 1Ca(HCO_3)_2 + 1Mg(OH)_2$	How many moles of $Mg(HCO_3)_2$?
37.5 g K	12.	$3K + 1AlCl_3 \rightarrow 3KCl + 1Al$	How many moles of $AlCl_3$?
67.5 g H_2SO_4	13.	$3H_2SO_4 + 1Ca_3(PO_4)_2 \rightarrow 2H_3PO_4 + 3CaSO_4$	How many grams of $Ca_3(PO_4)_2$?
2.25 g Cl_2	14.	$1Cl_2 + 1MgBr_2 \rightarrow 1MgCl_2 + 1Br_2$	How many grams of $MgBr_2$?
5.22 mol H_2CrO_4	15.	$1H_2S + 1PbCrO_4 \rightarrow 1H_2CrO_4 + 1PbS$	How many moles of PbS?
2.55 mol Ag_2S	16.	$1H_2S + 2Ag \rightarrow 1Ag_2S + 1H_2$	How many moles of H_2?
2.50 mol CO_2	17.	$2C_8H_{18} + 25O_2 \rightarrow 16CO_2 + 18H_2O$	How many grams of H_2O?
0.125 mol H_2O	18.	$1C_{25}H_{52} + 38O_2 \rightarrow 25CO_2 + 26H_2O$	How many grams of CO_2?
67.5 g H_2	19.	$1Fe + 2H_2O \rightarrow 1Fe(OH)_2 + 1H_2$	How many grams of Fe?
101 g $KClO_3$	20.	$2KClO_3 \rightarrow 2KCl + 3O_2$	How many moles of KCl?